1. The atomic symbol of silver is ————
(a) Si
(b) S
(c) Au
(d) Ag
Ans: (d)
Solution: The atomic symbol of silver is Ag
2. A box contains some identical red colour balls labelled as A each weighing 2 g. Another box contains identical blue colored balls, labelled as B, each weighing 5 g. In the combinations AB, AB ₂ , A ₂ B and A ₂ B ₃ which is applicable?
(a) Law of Definite proportion
(b) Law of multiple proportion
(c) Law of conservation of mass
(d) None of the above
Ans: (b)
Solution: According to the law of multiple proportion if an element forms more than one compound with another element for a given mass of an element, masses of other elements are in the ratio of the small whole number.
3. What is the value of Avogadro's number?
(a) 6.02 x 10 ⁻²³
(b) 6.02 x 10 ²³
(c) 6.02 x 10 ⁻²²
(d) 6.02 x 10 ²²
Ans: (b)
Solution: The value of Avogadro's number is 6.02 x 10 ²³
4. What is the chemical formula of sodium carbonate?
(a) Na ₂ CO ₃
(b) NaHCO ₃
(c) NaCO ₃
(d) Na ₂ HCO ₃
Ans: (a)
Solution: The chemical formula of sodium carbonate is Na ₂ CO ₃
5. Which of the following is the correct pair of atom and its atomic symbol?
(a) Sulphur – Su
(b) Potassium – P
(c) Phosphorus -P

(d) Sodium- S
Ans: (c)
Solution: The correct pair of atom and its atomic symbol is Phosphorus -P
6. How many moles are present in 40 g of He?
(a) 5 moles
(b) 20 moles
(c) 6 moles
(d) 10 moles
Ans: (d)
Solution: 1 mole of He = 4 gm. \Rightarrow 40 gm of He = 10 moles
7. Atomic mass of Chlorine is ——— (u)
(a) 34
(b) 34.5
(c) 35
(d) 35.5
Ans: (d)
Atomic mass of Chlorine is 35.5 u
8. 46 gm of sodium is equal to ———-
(a) 7 moles
(b) 1.5 moles
(c) 2 moles
(d) 1 mole
Ans: (c)
Solution: 1 mole of sodium = 23gm, 46 gm of sodium = 2 moles
9. The formula of Ammonium Sulphate is —————
(a) NH ₄ SO ₄
(b) NH ₄ SO ₂
(c) (NH ₄) ₂ SO ₄
(d) NH ₂ SO ₄
Ans: (c)
Solution: The formula of Ammonium Sulphate is (NH ₄) ₂ SO ₄ .
10. The atomic symbol of Iron is ————.
(a) I

(b) Fe

- (c) Ir
- (d) Au

Ans: (b)

Solution: The atomic symbol of Iron is Fe.